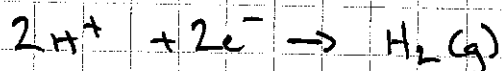
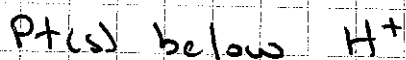
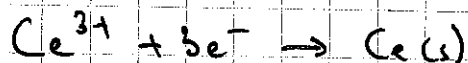
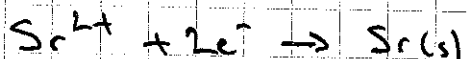
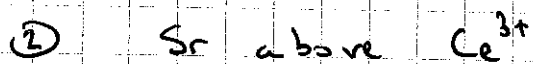
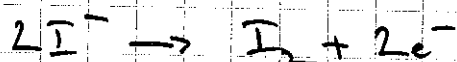
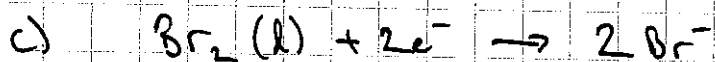
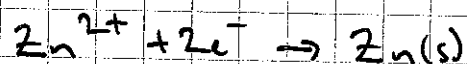
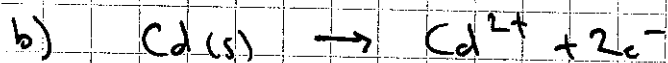
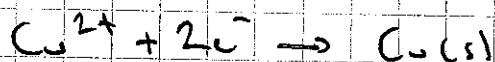


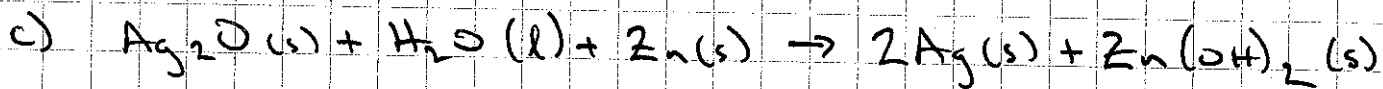
## Chemistry of Voltaic Cells



③ a) silver is gaining electrons (being reduced), so it is the cathode. That makes zinc the anode. Electrons flow from anode to cathode, so

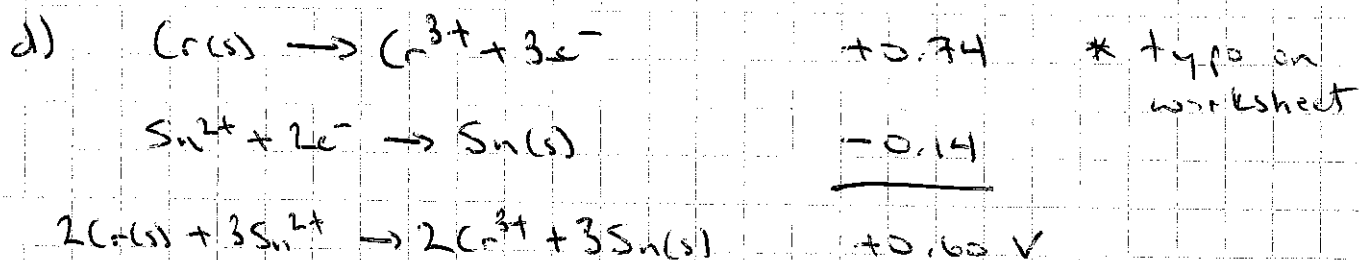
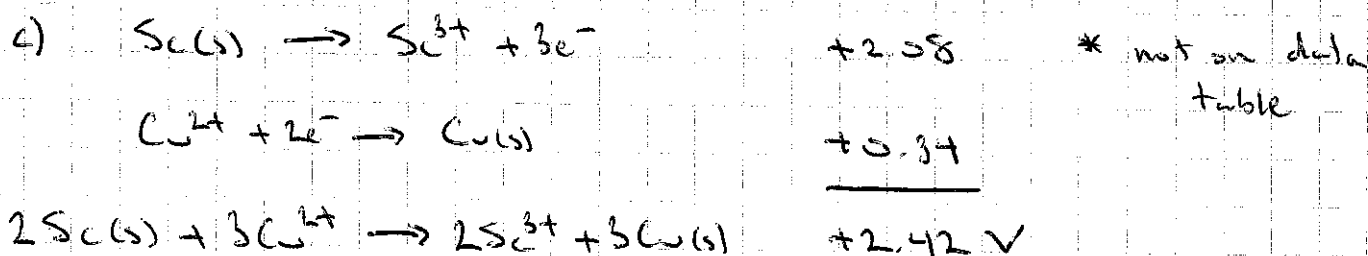
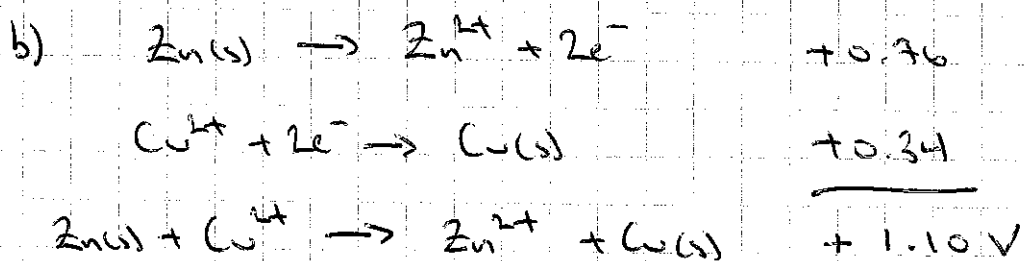
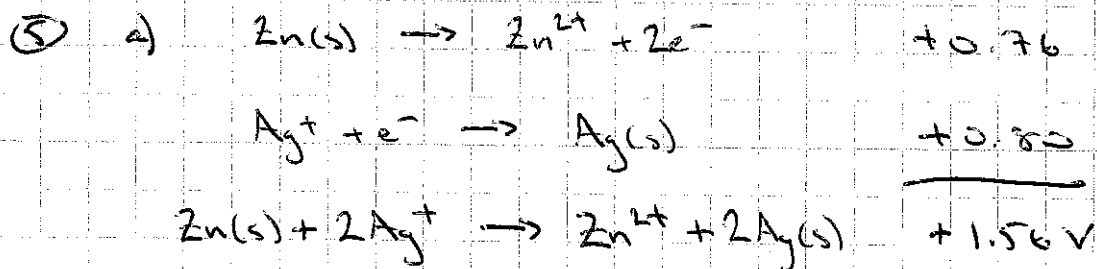
zinc to silver

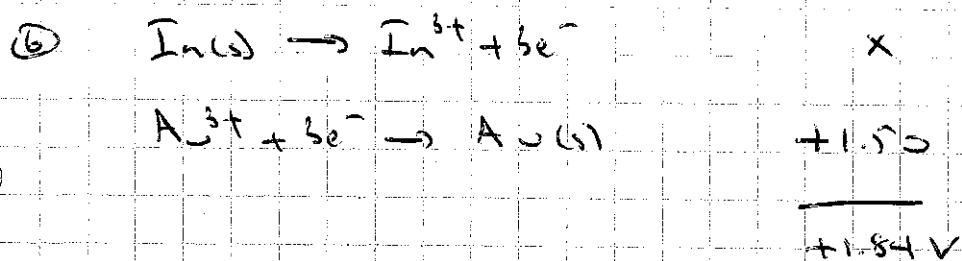
b) anode = zinc      cathode = silver



④ a) cathode

b) anode

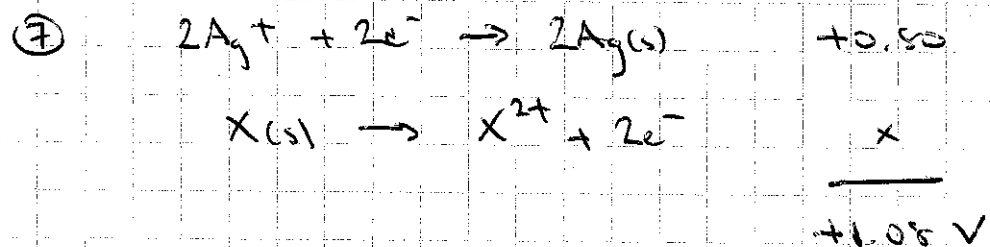




$$x + 1.50 = 1.84$$

$$x = 0.34 \text{ V} = \text{ox. potential}$$

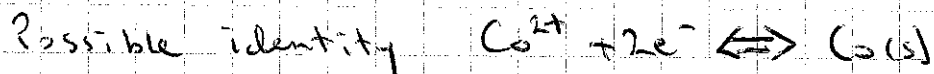
$$E_{\text{red}}^{\circ} = -0.34 \text{ V}$$



$$0.80 + x = 1.08$$

$$x = 0.28 \text{ V}$$

$$E_{\text{red}}^{\circ} = -0.28 \text{ V}$$



②

anode

